



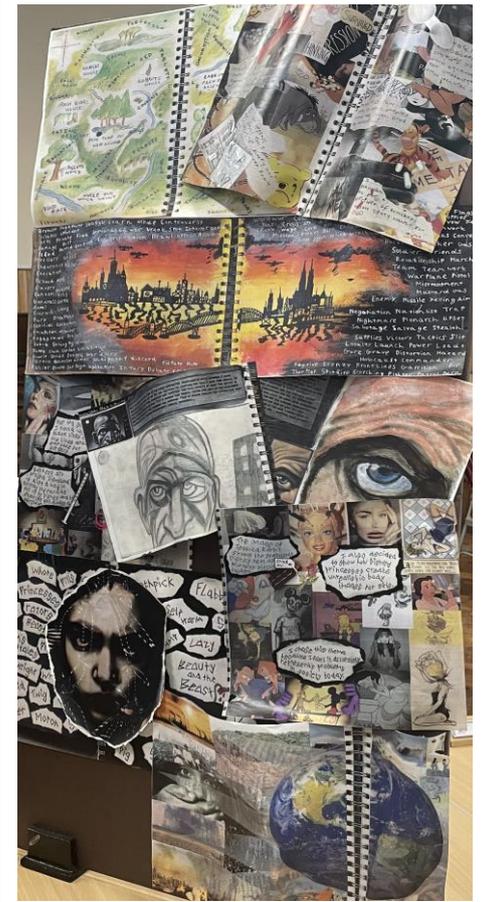
EBOR ACADEMY



Headteacher's update:

Last week we shared the joy of being able to open our doors and invite parents and pupils into school, in person, for our year 9 Options' Evening once again. It was wonderful to see so many of you here; I thoroughly enjoyed meeting parents and having the opportunity to talk together about our young people's aspirations as they began considering their future goals.

The subject 'stalls' on the evening exhibited an array of activities to be enjoyed during GCSE studies, and their incredible display of pupils' work, past and present, served to remind us of just how talented the young people of Filey town really are. I am immensely proud to be working with such gifted and talented individuals.



I am hopeful that the evening helped to provide a better insight into the exciting GCSE courses our year 9 pupils are about to embark on but completely understand that the transition into KS4 for some children can seem a little daunting as they make very 'grown up' decisions about their option choices. If you feel your child and/or you would like any further information about the transition into GCSE, please do contact us; we are more than happy to help in any way we can.



The week ahead brings us ever closer to our current year 11 students' exams. As I've been walking about lessons recently, I have been impressed by the quality of revision taking place. It's great to see some truly excellent examples of revision strategies being applied so assiduously by students. The examples you see here are from two wonderful young ladies who have found the creation of these revision maps a helpful revision tool. Encouraging retrieval exercises such as these is extremely beneficial. Why not ask your son or daughter about the things they have been doing to help them revise so far?

"I hope you all have a fabulous week and that the more clement weather we are experiencing enables you to enjoy exploring the outside with family and friends more readily.

Warmest wishes,
Miss Britton.

Revision Page

Reactivity Series

How well a metal reacts

The reactivity series lists metals in order of their reactivity towards other substances. For metals their reactivity is determined by how easily they lose electrons, forming positive ions. The higher up the reactivity is the more easily it reacts.

When metals react with water or acid they lose electrons, and form positive ions. The more higher a metal is easier it reacts with water or acids.

most reactive to least reactive = reactivity series

Acids and Bases

pH Scale goes from 0 to 14 and is measured by how acidic or alkaline solution is. Lower the pH scale is more acidic it is. Higher the more alkaline it is. Neutral = pH 7

Electrolysis

Electrolysis is used an electrical current. The positive ions will move to cathode. Negative ions move to anion (-ve) and lose electrons (+ve)

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Revision Questions

- Acids
- pH value acidic, Neutral or Alkali
- N=7 Acid = 2, Alkali = 9
- an equation for acids, Bases
- Acid + Base → Salt + water
- What is a Strong acid
- a Strong acid is an acid that is strongly

"I have been impressed by the quality of revision taking place."

Revision

Acids and bases

Measuring the pH:

- Wide range indicators = substances that gradually change colour depending on pH changes. e.g. universal indicator.

Acids/bases neutralisation

- When acids dissolve in water, they form solutions with a pH less than 7.
- Acids form H⁺ ions in water
- bases have a pH greater than 7
- Alkalis are bases which dissolve in water to form solutions with a pH greater than 7
- Alkalis form OH⁻ ions in water
- acid + base → salt + water
- H⁺ + OH⁻ → H₂O

Metal oxides/hydroxides

- bases
- react with acids in neutralisation reactions to form a salt and water
- the salt depends on the acid and the metal ion in the oxide/hydroxide

For example:

Sulphuric acid + potassium hydroxide → potassium sulphate + water

$$H_2SO_4 + 2KOH \rightarrow K_2SO_4 + 2H_2O$$

Acids and metal carbonates produce carbon dioxide

- Metal carbonates are bases. react with acids to produce salt, water and carbon dioxide

hydrochloric acid + sodium carbonate → sodium chloride + water + carbon dioxide

The reactivity series

most reactive → least reactive

- Potassium
- Sodium
- Lithium
- Calcium
- Magnesium
- Carbon
- Zinc
- Iron
- Hydrogen
- Copper

Cl, Br, I, Ag, Au

"I'm very much looking forward to sharing more of the wonderful work that goes on in our classrooms here over the coming weeks."